

# percent composition worksheet 2 answer key

**Percent composition worksheet 2 answer key** is a crucial educational resource for students studying chemistry. Understanding percent composition is fundamental in stoichiometry, allowing students to determine the relative amounts of different elements within a compound. This article explores the concept of percent composition, provides example problems, and outlines how to solve them, ultimately guiding you to the answer key for worksheet 2.

## What is Percent Composition?

Percent composition is the percentage by mass of each element in a compound. It provides insight into the chemical makeup of substances, which is vital for various applications in chemistry, such as calculating molar masses and performing reactions.

The formula for calculating the percent composition of an element in a compound is:

$$\text{Percent Composition} = \left( \frac{\text{Mass of the element in 1 mole of compound}}{\text{Molar mass of the compound}} \right) \times 100$$

## Why is Percent Composition Important?

Understanding percent composition is essential for several reasons:

- It helps in determining the empirical and molecular formulas of compounds.
- It aids in stoichiometric calculations in chemical reactions.
- It allows chemists to assess the purity of substances.
- It is fundamental in the field of materials science and engineering.

# Example Problems and Solutions

To illustrate the concept of percent composition, let's work through a few example problems.

## Example Problem 1: Water (H<sub>2</sub>O)

1. Step 1: Determine the molar mass of water.

- Hydrogen (H):  $1.01 \text{ g/mol} \times 2 = 2.02 \text{ g/mol}$
- Oxygen (O):  $16.00 \text{ g/mol} \times 1 = 16.00 \text{ g/mol}$
- Total molar mass of H<sub>2</sub>O =  $2.02 \text{ g/mol} + 16.00 \text{ g/mol} = 18.02 \text{ g/mol}$

2. Step 2: Calculate the percent composition.

- Percent composition of H:

$$\left[ \left( \frac{2.02 \text{ g/mol}}{18.02 \text{ g/mol}} \right) \times 100 = 11.21\% \right]$$

- Percent composition of O:

$$\left[ \left( \frac{16.00 \text{ g/mol}}{18.02 \text{ g/mol}} \right) \times 100 = 88.79\% \right]$$

Thus, the percent composition of water is approximately 11.21% hydrogen and 88.79% oxygen.

## Example Problem 2: Calcium Carbonate (CaCO<sub>3</sub>)

1. Step 1: Find the molar mass of calcium carbonate.

- Calcium (Ca):  $40.08 \text{ g/mol}$
- Carbon (C):  $12.01 \text{ g/mol} \times 1 = 12.01 \text{ g/mol}$
- Oxygen (O):  $16.00 \text{ g/mol} \times 3 = 48.00 \text{ g/mol}$
- Total molar mass of CaCO<sub>3</sub> =  $40.08 \text{ g/mol} + 12.01 \text{ g/mol} + 48.00 \text{ g/mol} = 100.09 \text{ g/mol}$

2. Step 2: Calculate the percent composition.

- Percent composition of Ca:

$$\left[ \left( \frac{40.08 \text{ g/mol}}{100.09 \text{ g/mol}} \right) \times 100 = 40.00\% \right]$$

- Percent composition of C:

$$\left[ \left( \frac{12.01 \text{ g/mol}}{100.09 \text{ g/mol}} \right) \times 100 = 11.98\% \right]$$

\]

- Percent composition of O:

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$$\left( \frac{48.00 \text{ g/mol}}{100.09 \text{ g/mol}} \right) \times 100 = 47.99\%$$

\]

Therefore, the percent composition of calcium carbonate is approximately 40.00% calcium, 11.98% carbon, and 47.99% oxygen.

## Percent Composition Worksheet 2

In a typical classroom setting, percent composition worksheets serve as practice tools for students. Worksheet 2 often contains various compounds for students to analyze.

Here are some compounds you might find in Worksheet 2:

1. Sodium Chloride (NaCl)
2. Sulfuric Acid (H<sub>2</sub>SO<sub>4</sub>)
3. Glucose (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>)

## Answer Key for Percent Composition Worksheet 2

Now that we have reviewed the concepts and example problems, let's provide a concise answer key for common compounds found in a typical percent composition worksheet.

Compound: Sodium Chloride (NaCl)

- Molar Mass:
- Sodium (Na): 22.99 g/mol
- Chlorine (Cl): 35.45 g/mol
- Total = 58.44 g/mol

- Percent Composition:

- Na:

\[

$$\left( \frac{22.99 \text{ g/mol}}{58.44 \text{ g/mol}} \right) \times 100 \approx 39.33\%$$

\]

- Cl:

\[

$$\left( \frac{35.45 \text{ g/mol}}{58.44 \text{ g/mol}} \right) \times 100 \approx 60.67\%$$

\]

Compound: Sulfuric Acid ( $\text{H}_2\text{SO}_4$ )

- Molar Mass:

- Hydrogen (H):  $1.01 \text{ g/mol} \times 2 = 2.02 \text{ g/mol}$

- Sulfur (S):  $32.07 \text{ g/mol}$

- Oxygen (O):  $16.00 \text{ g/mol} \times 4 = 64.00 \text{ g/mol}$

- Total =  $98.09 \text{ g/mol}$

- Percent Composition:

- H:

$$\left[ \left( \frac{2.02 \text{ g/mol}}{98.09 \text{ g/mol}} \right) \times 100 \right] \approx 2.06\%$$

- S:

$$\left[ \left( \frac{32.07 \text{ g/mol}}{98.09 \text{ g/mol}} \right) \times 100 \right] \approx 32.66\%$$

- O:

$$\left[ \left( \frac{64.00 \text{ g/mol}}{98.09 \text{ g/mol}} \right) \times 100 \right] \approx 65.28\%$$

Compound: Glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ )

- Molar Mass:

- Carbon (C):  $12.01 \text{ g/mol} \times 6 = 72.06 \text{ g/mol}$

- Hydrogen (H):  $1.01 \text{ g/mol} \times 12 = 12.12 \text{ g/mol}$

- Oxygen (O):  $16.00 \text{ g/mol} \times 6 = 96.00 \text{ g/mol}$

- Total =  $180.18 \text{ g/mol}$

- Percent Composition:

- C:

$$\left[ \left( \frac{72.06 \text{ g/mol}}{180.18 \text{ g/mol}} \right) \times 100 \right] \approx 40.00\%$$

- H:

$$\left[ \left( \frac{12.12 \text{ g/mol}}{180.18 \text{ g/mol}} \right) \times 100 \right] \approx 6.73\%$$

- O:

$$\left[ \left( \frac{96.00 \text{ g/mol}}{180.18 \text{ g/mol}} \right) \times 100 \right] \approx 53.27\%$$

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# Conclusion

Understanding percent composition is a fundamental skill in chemistry that enhances students' analytical and problem-solving abilities. The percent composition worksheet 2 answer key serves as a valuable tool for students to check their understanding and accuracy in calculating the percent composition of various compounds. Mastering this concept will not only aid in academic pursuits but also provide a solid foundation for future studies in chemistry and related fields.

## Frequently Asked Questions

### What is a percent composition worksheet used for?

A percent composition worksheet is used to calculate the percentage by mass of each element in a compound.

### Where can I find the answer key for percent composition worksheet 2?

The answer key for percent composition worksheet 2 can typically be found in the teacher's guide or educational resource websites that provide chemistry worksheets.

### How do you calculate percent composition?

To calculate percent composition, divide the mass of the element in the compound by the total mass of the compound, then multiply by 100.

### What are common compounds used in percent composition worksheets?

Common compounds include water ( $\text{H}_2\text{O}$ ), sodium chloride ( $\text{NaCl}$ ), and glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ).

### Why is understanding percent composition important in chemistry?

Understanding percent composition is important for stoichiometry, determining empirical formulas, and analyzing chemical reactions.

### Are there online resources for percent composition

## **practice?**

Yes, there are many online platforms that offer practice problems and worksheets, including Khan Academy, ChemCollective, and educational websites.

## **What skills do students develop by completing percent composition worksheets?**

Students develop skills in critical thinking, problem-solving, and quantitative analysis by completing percent composition worksheets.

## **Can percent composition be applied in real-world scenarios?**

Yes, percent composition is used in various fields such as pharmaceuticals, nutrition, and environmental science to analyze mixtures and compounds.

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