

# oxidation number worksheet with answers

Oxidation number worksheet with answers is an essential tool for students and educators in the field of chemistry. Understanding oxidation numbers is crucial for mastering concepts related to oxidation-reduction (redox) reactions, balancing chemical equations, and comprehending the behavior of various chemical species. This article will delve into the concept of oxidation numbers, provide a worksheet with exercises, and supply detailed answers for each question, ensuring a comprehensive understanding of this important subject.

## Understanding Oxidation Numbers

Oxidation numbers, also known as oxidation states, are assigned to atoms in a compound or ion to reflect the number of electrons that an atom gains, loses, or shares when forming a chemical bond. The concept is vital for determining how electrons are transferred in redox reactions.

## Basic Rules for Assigning Oxidation Numbers

To accurately assign oxidation numbers, it is essential to follow several fundamental rules:

1. **Elemental State:** The oxidation number of an atom in its elemental form (e.g.,  $O_2$ ,  $N_2$ , Fe) is always zero.
2. **Monatomic Ions:** The oxidation number of a monatomic ion is equal to the charge of the ion (e.g.,  $Na^+$  has an oxidation number of +1,  $Cl^-$  has an oxidation number of -1).
3. **Hydrogen:** Hydrogen usually has an oxidation number of +1 when bonded to nonmetals and -1 when bonded to metals.
4. **Oxygen:** Oxygen typically has an oxidation number of -2 in most compounds, with exceptions such as peroxides, where it is -1.
5. **Alkali Metals:** Alkali metals (Group 1) have an oxidation number of +1 in all their compounds.
6. **Alkaline Earth Metals:** Alkaline earth metals (Group 2) have an oxidation number of +2 in their compounds.
7. **Sum of Oxidation Numbers:** The sum of the oxidation numbers in a neutral compound must equal zero, while in a polyatomic ion, it must equal the charge of the ion.

8. Fluorine: Fluorine always has an oxidation number of -1 in compounds.

## Importance of Oxidation Numbers

Understanding oxidation numbers is crucial for several reasons:

- Identifying Redox Reactions: Oxidation numbers help identify the species being oxidized and reduced in a reaction.
- Balancing Chemical Equations: They provide a systematic way to balance equations involving redox processes.
- Predicting Reactivity: Knowing the oxidation states can help predict the reactivity of certain elements and compounds.

## Oxidation Number Worksheet

Below is a worksheet designed to help reinforce the concept of oxidation numbers. Each question requires the student to determine the oxidation number of the specified element in each compound or ion.

Worksheet: Oxidation Numbers

1. Determine the oxidation number of sulfur in  $\text{H}_2\text{SO}_4$ .
2. Calculate the oxidation number of nitrogen in  $\text{NH}_3$ .
3. What is the oxidation number of carbon in  $\text{CO}_2$ ?
4. Find the oxidation number of chlorine in  $\text{ClO}_3^-$ .
5. Determine the oxidation state of copper in  $\text{CuSO}_4$ .
6. What is the oxidation number of iron in  $\text{Fe}_2\text{O}_3$ ?
7. Calculate the oxidation number of manganese in  $\text{KMnO}_4$ .
8. What is the oxidation number of phosphorus in  $\text{H}_3\text{PO}_4$ ?
9. Determine the oxidation state of chromium in  $\text{Cr}_2\text{O}_7^{2-}$ .
10. Find the oxidation number of silver in  $\text{AgNO}_3$ .

## Answers to the Oxidation Number Worksheet

Now that the worksheet has been provided, let's go through each question and present the correct answers along with explanations.

## 1. Sulfur in $\text{H}_2\text{SO}_4$

- Answer: +6

- Explanation: The sum of oxidation numbers must equal zero. Hydrogen is +1 ( $2 \times +1 = +2$ ), and oxygen is -2 ( $4 \times -2 = -8$ ). Thus, for sulfur:  $+2 + x - 8 = 0 \rightarrow x = +6$ .

## 2. Nitrogen in $\text{NH}_3$

- Answer: -3

- Explanation: Hydrogen is +1 ( $3 \times +1 = +3$ ). Thus, for nitrogen:  $+3 + x = 0 \rightarrow x = -3$ .

## 3. Carbon in $\text{CO}_2$

- Answer: +4

- Explanation: Oxygen is -2 ( $2 \times -2 = -4$ ). Therefore, for carbon:  $x - 4 = 0 \rightarrow x = +4$ .

## 4. Chlorine in $\text{ClO}_3^-$

- Answer: +5

- Explanation: Oxygen is -2 ( $3 \times -2 = -6$ ). The ion has a charge of -1. For chlorine:  $x - 6 = -1 \rightarrow x = +5$ .

## 5. Copper in $\text{CuSO}_4$

- Answer: +2

- Explanation: Sulfate ( $\text{SO}_4^{2-}$ ) has a charge of -2. For copper:  $x - 2 = 0 \rightarrow x = +2$ .

## 6. Iron in $\text{Fe}_2\text{O}_3$

- Answer: +3

- Explanation: Oxygen is -2 ( $3 \times -2 = -6$ ). For iron:  $2x - 6 = 0 \rightarrow 2x = 6 \rightarrow x = +3$ .

## 7. Manganese in $\text{KMnO}_4$

- Answer: +7

- Explanation: Potassium (K) is +1, and oxygen is -2 ( $4 \times -2 = -8$ ). For manganese:  $+1 + x - 8 = 0 \rightarrow x = +7$ .

## 8. Phosphorus in $\text{H}_3\text{PO}_4$

- Answer: +5

- Explanation: Hydrogen is +1 ( $3 \times +1 = +3$ ), and oxygen is -2 ( $4 \times -2 = -8$ ). For phosphorus:  $+3 + x - 8 = 0 \rightarrow x = +5$ .

## 9. Chromium in $\text{Cr}_2\text{O}_7^{2-}$

- Answer: +6

- Explanation: Oxygen is -2 ( $7 \times -2 = -14$ ). The total charge is -2. For chromium:  $2x - 14 = -2 \rightarrow 2x = 12 \rightarrow x = +6$ .

## 10. Silver in $\text{AgNO}_3$

- Answer: +1

- Explanation: Nitrate ( $\text{NO}_3^-$ ) has a charge of -1. For silver:  $x - 1 = 0 \rightarrow x = +1$ .

## Conclusion

The oxidation number worksheet with answers serves as a vital resource for students seeking to enhance their understanding of oxidation states. Through practice, students can improve their ability to identify and calculate oxidation numbers, which is an essential skill in chemistry. By mastering these concepts, learners will find themselves better equipped to tackle more complex topics, such as redox reactions and balancing chemical equations. Regular practice with worksheets like the one provided can lead to increased confidence and proficiency in chemistry.

## Frequently Asked Questions

### What is an oxidation number?

An oxidation number is a value assigned to an element in a compound that represents the number of electrons lost or gained by that element during a chemical reaction.

### How do you determine the oxidation number of an element in a compound?

To determine the oxidation number, identify the common oxidation states of the elements, apply the rules for oxidation numbers, and ensure the sum of oxidation numbers equals the overall charge of the

compound.

## **What are the general rules for assigning oxidation numbers?**

Some general rules include: 1) The oxidation number of an element in its standard state is 0. 2) The oxidation number of a monoatomic ion equals its charge. 3) Oxygen typically has an oxidation number of -2, and hydrogen is usually +1.

## **Can you provide an example of a compound and its oxidation numbers?**

In water ( $\text{H}_2\text{O}$ ), the oxidation number of hydrogen is +1 and oxygen is -2. Thus, the overall charge is neutral, as  $(2 \times +1) + (-2) = 0$ .

## **What is the oxidation number of sulfur in sulfuric acid ( $\text{H}_2\text{SO}_4$ )?**

In sulfuric acid ( $\text{H}_2\text{SO}_4$ ), the oxidation number of sulfur is +6.

## **How can oxidation numbers help in balancing redox reactions?**

Oxidation numbers help identify which elements are oxidized and reduced in a reaction, allowing for the systematic balancing of redox equations by ensuring that the total increase and decrease of oxidation numbers are equal.

## **What is the oxidation state of carbon in carbon dioxide ( $\text{CO}_2$ )?**

In carbon dioxide ( $\text{CO}_2$ ), the oxidation state of carbon is +4.

## **Why is it important to understand oxidation numbers in chemistry?**

Understanding oxidation numbers is crucial for predicting reaction outcomes, determining electron transfer in redox reactions, and comprehending the behavior of various chemical species.

## **Are there exceptions to the oxidation number rules?**

Yes, there are exceptions to the rules, such as in peroxides where oxygen has an oxidation number of -1, and in compounds with metals where the oxidation states may vary based on the specific compound.

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